WORKSHEET#7 examples, CHEM001
(Chapter 7)

Name: _____ **Date:** _____

1. **MOLAR MASS:** What is the molar mass of $MgSO_4 * 6H_2O$?

2. **PERCENT OF ONE ELEMENT:** What percent of the molar mass of H_2SO_4 is made up by oxygen? How many grams of oxygen are contained in 12 grams of H_2SO_4?

3. **CONVERSION MAP:** How many atoms of hydrogen are contained in 17 g of C_4H_{10}?

WORKSHEET#7, CHEM001
(Chapter 7)

Name: _____ **Date:** _____

1. **MOLAR MASS:** What is the molar mass of $BaCl_2 * 2H_2O$?

2. **PERCENT OF ONE ELEMENT:** What percent of the molar mass of $CuSO_4$ is made up by oxygen? How many grams of oxygen are contained in 54 grams of $CuSO_4$?

3. **CONVERSION MAP:** (7-22) How many atoms of hydrogen are contained in 57 g of $C_6H_5ONH_2$?

4. **EMPIRICAL/MOLECULAR FORMULA:** (7-48) A compound is composed of 54.5% C, 9.2% H, and 36.3% O. It has a molar mass of 88.11 g/mol. Determine the empirical and molecular formulas for this substance.

5. **EMPIRICAL/MOLECULAR FORMULA:** (7-44) A compound is composed of 29.73% C, 4.16% H, 13.20% O, and 52.91% S. It has a molar mass of 242.4 g/mol. Determine the empirical and molecular formulas for this substance.